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Assignment: Combined Gas Law

1. The initial temperature of a 1.00 L sample of argon is 20.0°C. The pressure is decreased from 720mmHg to 360mmHg and the volume increases to 2.14 L. What is the final temperature of the argon gas?
2. A sample of nitrogen gas occupies a volume of 2.00 L at 756mmHg and 0.00°C. The volume increases by 2.00 L and the temperature decreases to 137 K. What is the final pressure exerted on the gas?
3. A 20.0L container is filled with helium at a pressure of 150atm and a temperature of 30.0°C. How many 5.0 L balloons can be filled when the temperature is 22°C and the atmospheric pressure is 755mmHg?